

$$(2S+1)L_J$$

$$S = \frac{1}{2}$$

$$J = L \oplus S$$

$$\lambda_{3p-2s} = 656.27 \text{ nm}$$

$$\lambda_{3d-2p} = 656.28 \text{ nm}$$

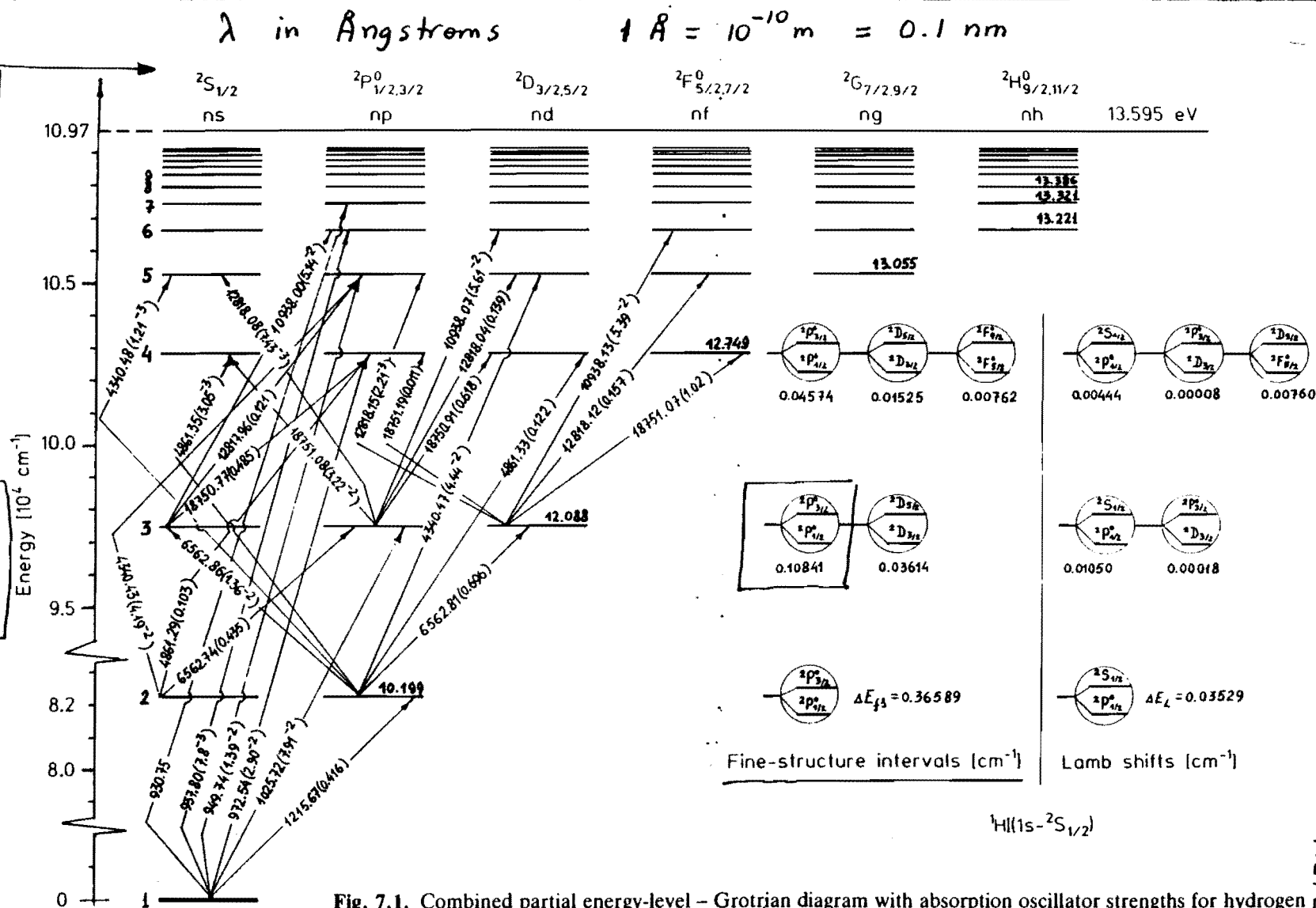


Fig. 7.1. Combined partial energy-level - Grotjan diagram with absorption oscillator strengths for hydrogen